

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                                    |                   |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair        | 3. weak acid      |
| 4. hydronium      | 5. hydroxide ion ( $\text{OH}^-$ ) | 6. autoionization |
|                   | 7. weak base                       | 8. pH             |

## Key Equations

- |  |   |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3)                                  | 4. $\text{pOH} = -\log[\text{OH}^-]$ (14.4)   |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (strong hydroxide base/weak hydronium acid)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (the autoionization of water which produces an acidic hydronium ion and a basic hydroxide ion)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (weak hydroxide base/strong hydronium acid)
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (strong hydroxide base/strong hydronium acid)
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (the autoionization of water which produces a weak hydronium acid and a weak hydroxide base)
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (strong hydroxide base/weak hydronium acid)

7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
8.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
10.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
11.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
12.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
13.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
14.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
15.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
16.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

# 19 2 Hydrogen Ions And Acidity Section Review

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## **Table of Contents 19 2 Hydrogen Ions And Acidity Section Review**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Section Review User Reviews and Ratings
  - 19 2 Hydrogen Ions And Acidity Section Review and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Free and Paid eBooks
  - 19 2 Hydrogen Ions And Acidity Section Review Public Domain eBooks

- 19 2 Hydrogen Ions And Acidity Section Review eBook Subscription Services
- 19 2 Hydrogen Ions And Acidity Section Review Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Section Review eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Section Review Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Section Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review
  - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review
  - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review
  - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review
  - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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