

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (B) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (D) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$ (E) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (F) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. In a 0.1 mol/L acetic acid solution, the concentration of acetate ions is 0.01 mol/L. What is the concentration of acetic acid? (correct)

(A) 0.01 mol/L (B) 0.09 mol/L

(C) 0.01 mol/L (D) 0.09 mol/L (E) 0.01 mol/L (F) 0.09 mol/L

3. In a 0.1 mol/L acetic acid solution, the concentration of acetate ions is 0.01 mol/L. What is the concentration of acetic acid? (correct)

(A) 0.01 mol/L (B) 0.09 mol/L (C) 0.01 mol/L (D) 0.09 mol/L

4. Which of the following is a weak acid? (correct)



5. A solution of acetic acid and its conjugate base is prepared by mixing 100 mL of 0.1 mol/L acetic acid and 100 mL of 0.1 mol/L sodium acetate. What is the pH of the solution? (correct)

(A) 4.75 (B) 4.75 (C) 4.75 (D) 4.75 (E) 4.75

(F) 4.75 (G) 4.75 (H) 4.75 (I) 4.75 (J) 4.75

6. In a 0.1 mol/L acetic acid solution, the concentration of acetate ions is 0.01 mol/L. What is the concentration of acetic acid? (correct)



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**Thirumalaisamy P.
Velavancorresponding**

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