

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
2. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/weak base)
3. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
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9. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
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