

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base indicator has a  $\text{pK}_a$  of 4.5. At what pH will the indicator be 10% in its protonated form and 90% in its deprotonated form? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 3.5$       (C)  $\text{pH} = 5.5$       (D)  $\text{pH} = 4.0$       (E)  $\text{pH} = 5.0$

3. An acid-base indicator has a  $\text{pK}_a$  of 4.5. At what pH will the indicator be 10% in its protonated form and 90% in its deprotonated form? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 3.5$       (C)  $\text{pH} = 5.5$       (D)  $\text{pH} = 4.0$       (E)  $\text{pH} = 5.0$

4. Which of the following is a buffer solution? (correct)



(C)  $\text{HCl}$       (D)  $\text{NaOH}$       (E)  $\text{HCl}$  and  $\text{NaOH}$

5. A buffer solution contains 0.1 M of a weak acid and 0.1 M of its conjugate base. The  $\text{pK}_a$  of the weak acid is 4.5. What is the pH of the buffer solution? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 3.5$       (C)  $\text{pH} = 5.5$       (D)  $\text{pH} = 4.0$       (E)  $\text{pH} = 5.0$

(F)  $\text{pH} = 4.5$       (G)  $\text{pH} = 3.5$       (H)  $\text{pH} = 5.5$       (I)  $\text{pH} = 4.0$       (J)  $\text{pH} = 5.0$

6. An acid-base indicator has a  $\text{pK}_a$  of 4.5. At what pH will the indicator be 10% in its protonated form and 90% in its deprotonated form? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 3.5$       (C)  $\text{pH} = 5.5$       (D)  $\text{pH} = 4.0$       (E)  $\text{pH} = 5.0$

# Multiple Choice Stoichiometry Test

**Camilla Rothe**

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