

### 18.4 Molecular Structure: Mass

#### Multiple Choice:

1. The molar masses of quantities in chemical equations are based on \_\_\_\_\_  
 (A) percent composition (B) percent yield  
 (C) molar formula (D) molar structural analysis
2. Percent of hydrogen in  $\text{C}_2\text{H}_6\text{O}$  is \_\_\_\_\_ of the total mass of the compound?

(A) 15.1% (B) 12.5%

(C) 15.1% (D) 12.5% (E) 15.1% (F) percent of the atoms

3. In a gas the other reactions represent oxygen mixed with other elements. If 1.1 g of gas was produced by the reaction, and only produced 1.1 g of gas, then the gas percent of the gas is \_\_\_\_\_?

(A) 15.1% (B) 12.5% (C) 15.1% (D) 12.5%

(E) 15.1% (F) 12.5% (G) 15.1% (H) 12.5%

4. In the reaction  $\text{C}_2\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{C}_2\text{H}_4\text{O} + \text{H}_2\text{O}$ , what is the percent of oxygen in the product?

(A) 15.1% (B) 12.5% (C) 15.1% (D) 12.5%

5. What is the percent of oxygen in the following equation?



(A) 15.1% (B) 12.5% (C) 15.1% (D) 12.5%

6. A mixture of water and H<sub>2</sub> gas is placed in a container. The mixture is then heated. If 1.1 g of water is placed and 1.1 g of H<sub>2</sub> gas is produced, what mass of H<sub>2</sub> gas is produced?



(A) 1.1 g (B) 1.1 g (C) 1.1 g (D) 1.1 g (E) 1.1 g

7. In a full experiment, structure-molecular weight of  $\text{C}_2\text{H}_6\text{O}$  is what percent of the total mass of  $\text{C}_2\text{H}_6\text{O}$ . After the reaction was completed, the total mass percent was 15.1% of the total mass of the reaction.



# Multiple Choice Stoichiometry Test

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## **Table of Contents Multiple Choice Stoichiometry Test**

1. Understanding the eBook Multiple Choice Stoichiometry Test
  - The Rise of Digital Reading Multiple Choice Stoichiometry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Stoichiometry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
  - Personalized Recommendations
  - Multiple Choice Stoichiometry Test User Reviews and Ratings
  - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
  - Multiple Choice Stoichiometry Test Public Domain eBooks
  - Multiple Choice Stoichiometry Test eBook Subscription Services
  - Multiple Choice Stoichiometry Test Budget-Friendly Options
6. Navigating Multiple Choice Stoichiometry Test eBook Formats

- ePub, PDF, MOBI, and More
  - Multiple Choice Stoichiometry Test Compatibility with Devices
  - Multiple Choice Stoichiometry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
    - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
    - Interactive Elements Multiple Choice Stoichiometry Test
  8. Staying Engaged with Multiple Choice Stoichiometry Test
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Multiple Choice Stoichiometry Test
  9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
    - Setting Reading Goals Multiple Choice Stoichiometry Test
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
    - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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