

CHEMISTRY I - 1st SEMESTER EXAM REVIEW

STRATEGY: Start by reading through your notes to refresh your memory on these topics. Then, use this review sheet as a starting point to identify the areas on which you need to spend more study time. For those areas, go back to homework assignments, quizzes, and reviews to practice more problems. I would also recommend going through the website links of tutorials and practice quizzes.

FORMAT:

- + Questions will be multiple-choice.
- + A formula bank will be provided in addition to any values that you might need (electronegativity, etc.), but you will NOT be given "formulas" for items listed in the VOCAB sections (average atomic mass, % error, etc.).

Atomic Structure – Ch. 3

1. Identify the scientists who made the following discoveries.
 - a. Atoms contain negative particles called electrons.
 - b. The mass of an electron is 9.11×10^{-28} g.
 - c. Atoms contain neutral particles called neutrons.
 - d. Atoms contain a dense, positive nucleus.
 - e. Atoms are indivisible and resemble billiard balls.
2. Describe the evolution of the atomic model from the billiard ball model to the electron cloud model.
3. Write the isotope symbol, including atomic number & mass number, for the following isotopes.
 - a. carbon-14 c. nickel-63
 - b. chromium-53 d. zirconium-92

4. Complete the table for the following isotopes.

Symbol	Zn			
Atomic #		20		
Mass #	65		74	40
# of protons			34	
# of neutrons		21		
# of electrons				18

5. Calculate the average atomic mass of copper if 69.17% of the copper atoms occurring in nature are ^{63}Cu and 30.83% are ^{65}Cu .

VOCAB: isotope average atomic mass
mole

Matter – Ch. 1

6. Classify the following substances as *solid, liquid, gas, or plasma* based on their properties.
 - a. flexible volume, high KE, particles can disperse freely.
 - b. flexible volume, very high KE, particles are charged.
 - c. fixed volume, very low KE, orderly particles.
 - d. fixed volume, low KE, particles can move past each other.
7. Compare and contrast a solution, colloid, and suspension.
8. Classify the following as element, compound, heterogeneous mixture, or solution.
 - a. graphite (carbon)
 - b. grape juice
 - c. table salt (NaCl)
 - d. pepper
9. Classify the following as *chemical or physical changes*.
 - a. cutting wire
 - b. ripening tomato
 - c. apple slices turning brown
 - d. compressing a gas
10. Classify the following properties as *physical or chemical*.
 - a. melts at 68.0°C
 - b. corrosive
 - c. reacts violently with water
 - d. decomposes in air
 - e. magnetic

VOCAB: kinetic molecular theory
law of definite composition
law of multiple proportions

Measurement – Ch. 2

11. In a lab, the average measured density for Pre-1982 pennies was 7.98 g/cm^3 . Given that the literature value for the density is 8.92 g/cm^3 , calculate the percent error.
12. How many sig figs are in the following numbers?
 - a. 2.35 c. 89.70
 - b. 34,000 d. 0.0052
13. Convert the following numbers into or out of scientific notation.
 - a. 548,000 c. 1.200×10^{-3}
 - b. 0.0000770 d. 9.25×10^7
14. Osmium is the densest element with a density of 22.57 g/cm^3 . Find the mass of a 56.2 cm^3 sample of osmium.
15. Perform the following SI prefix conversions.
 - a. $65.2 \text{ mm} = ? \text{ dm}$ c. $65,000 \text{ }\mu\text{L} = ? \text{ mL}$
 - b. $2.3 \text{ kg} = ? \text{ g}$ d. $0.502 \text{ km} = ? \text{ cm}$
16. How many milliliters are in a 2.0 quart jug of milk?
17. Mrs. J. spent last weekend grading lab reports. If she spent 5.5 min on each lab, how many hours did it take her to grade all 95 labs?
18. Calculate the density from the slope of a "Mass vs. Volume" graph.
19. Record the appropriate # of SigFigs when measuring.

VOCAB: accuracy vs. precision

Electrons in Atoms – Ch. 4 & 5

20. Calculate the wavelength if the frequency is $2.5 \times 10^8 \text{ Hz}$.
21. Find the energy of a photon if frequency is $7.31 \times 10^{14} \text{ Hz}$.
22. Describe how Bohr's model explains the bright lines (red, green, violet, violet) in the emission spectrum of hydrogen.
23. What is the primary difference between the modern model of the atom and Bohr's model?
24. Draw orbital diagrams for the following elements.

Symbol	Atomic #	Orbital Diagram
F		
V		

25. Explain why chromium's electron configuration is $[\text{Ar}] 4s^1 3d^5$ instead of the expected configuration of $[\text{Ar}] 4s^2 3d^4$.

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JR Anderson



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