

## Formal Charge

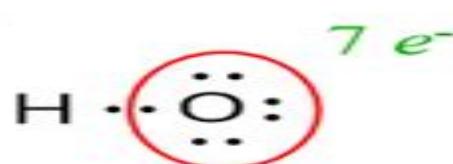
Oxygen is in group 6, so it has 6 valence  $e^-$ .



Valence  $e^-$  - 6

Assigned  $e^-$  - 6

Formal Charge = 0



Valence  $e^-$  - 6

Assigned  $e^-$  - 7

Formal Charge = -1

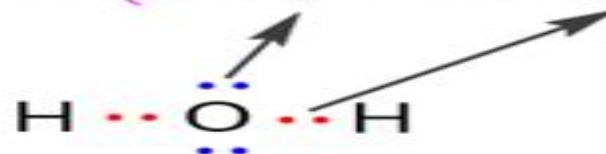
*Formal charge is the difference between the number of valence electrons and the sum of the nonbonding and half of the bonding electrons.*

$$\mathbf{FC} = \# \text{ valence } e^- -$$

$$- (\# \text{ nonbonding } e^- + 1/2 \# \text{ bonding } e^-)$$



$$\mathbf{FC} = 6 - (2 \times 2 + 0.5 \times 4) = 0$$



$$\mathbf{FC} = 6 - (3 \times 2 + 0.5 \times 2) = -1$$



## Formal Charges

**M Mosston**

**Formal Charges:**

## Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Formal Charges**

In a world inundated with displays and the cacophony of fast communication, the profound energy and mental resonance of verbal artistry usually diminish in to obscurity, eclipsed by the regular assault of noise and distractions. However, nestled within the musical pages of **Formal Charges**, a captivating function of literary splendor that impulses with natural thoughts, lies an remarkable trip waiting to be embarked upon. Penned with a virtuoso wordsmith, this mesmerizing opus instructions visitors on a mental odyssey, delicately exposing the latent potential and profound affect embedded within the complex internet of language. Within the heart-wrenching expanse with this evocative analysis, we will embark upon an introspective exploration of the book is central subjects, dissect their interesting publishing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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