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arrhenius model of bases a substance that contains a hydroxide group and dissociates to produce a hydroxide

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the OH^- basic solution one in which the H^+ is less than the OH^- alkaline solutions the H^+ of a basic solution is less than $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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NH_2 and S^{2-} are CH_4 NH_3 and HS^- respectively next we consider that acidity increases with positive charge on the molecule

thus ruling out that S^{2-} is the weakest base

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