

**KEY**

1 26. Which is the stronger base,  $\text{HCO}_3^-$  or  $\text{HCOO}^-$ ? (1 mark)  $\text{HCO}_3^-$

27. Classify each of the following as: a strong acid (SA), weak acid (WA), strong base (SB), weak base (WB) or a spectator ion (S). (10 marks)

10

- |                                |           |                             |           |
|--------------------------------|-----------|-----------------------------|-----------|
| a) $\text{F}^-$                | <u>WB</u> | f) $\text{Cl}^-$            | <u>S</u>  |
| b) $\text{HIO}_3$              | <u>WA</u> | g) $\text{NH}_3$            | <u>WB</u> |
| c) $\text{NO}_3^-$             | <u>S</u>  | h) $\text{O}^{2-}$          | <u>SB</u> |
| d) $\text{HClO}_4$             | <u>SA</u> | i) $\text{CH}_3\text{COOH}$ | <u>WA</u> |
| e) $\text{C}_2\text{O}_4^{2-}$ | <u>WB</u> | j) $\text{ClO}_4^-$         | <u>S</u>  |

28. What is the  $[\text{OH}^-]$  in a solution made by adding 0.060 moles of calcium oxide to 500.0 mL water? Be careful! (2 marks)

2

$$M = \frac{\text{mol}}{L} = \frac{0.060 \text{ mol CaO}}{0.500 \text{ L}} = 0.12 \text{ M} = [\text{CaO}]$$

$$0.12 \text{ CaO} \rightarrow \text{Ca}^{2+} + 0.12 \text{ O}^{2-} \Rightarrow \text{O}^{2-} + \text{H}_2\text{O} \rightarrow 2\text{OH}^-$$

Answer  $[\text{OH}^-] = 0.24 \text{ M}$

29. What is the  $[\text{H}_3\text{O}^+]$  in a solution made by adding 0.020 moles of nitric acid to 500.0 mL of water? (2 marks)

2

$$[\text{HNO}_3] = \frac{0.020 \text{ mol}}{0.500 \text{ L}} = 0.040 \text{ M}$$

SA so  $[\text{H}_3\text{O}^+] = [\text{acid}] = 0.040 \text{ M}$

Answer  $[\text{H}_3\text{O}^+] = 0.040 \text{ M}$

30. If 0.10 M  $\text{HSO}_3^-$  is mixed with 0.10 M  $\text{HC}_2\text{O}_4^-$ , which species will donate a proton? (1 mark)

1

Answer  $\text{HC}_2\text{O}_4^-$

31. If 0.10 M  $\text{HSO}_4^-$  is mixed with 0.10 M  $\text{HC}_6\text{H}_5\text{O}_2^-$ , which species will donate a proton? (1 mark)

1

Answer  $\text{HSO}_4^-$

$\frac{17}{17}$

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**Thomas Griffiths**



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## **Table of Contents 19 3 Acids And Bases Answer Key Chemistry**

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
  - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry

- Personalized Recommendations
  - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
  - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
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  - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
- ePub, PDF, MOBI, and More
  - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
  - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
  - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
  - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
- Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry

- Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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