
ACS CHEMISTRY FINAL EXAM STUDY GUIDE

- milli (m) (m has 3 lines, upside down it's a w (win) so negative=loss): 10^{-3}
- micro (inappropriate joke, i don't want Quizlet to ban me lol): 10^{-6}
- nano: 10^{-9}
- pico (p): 10^{-12}
- mono: one, single, alone
- di: two
- tri: three (tricycle)
- tetra: four (tetrapod) - dogs, cats, vertebrae animals
- penta: five (pentagon)
- hexa: 6 (hexagon)

- hepta: 7 (heptagon)
- octa: 8 (octopus, stop)
- nona: 9 (no = neign in German)
- deca: 10 (decade = 10 years)
- molar mass: the mass of one mole of a pure substance
- molar ratio: ratio of coefficients in a chemical equation
- limiting reactant: A reactant that is totally consumed during a chemical reaction, limits the extent of the reaction, and determines the amount of product.
- percent composition: the percent by mass of each element in a compound
- molarity: the number of moles of solute per liter of solution
- percent yield: $\text{actual yield/theoretical yield} \times 100$
- empirical formula: a chemical formula showing the ratio of elements in a compound rather than the total number of atoms
- molecular formula: A chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms.
- gas laws: describe how the behavior of gases is affected by pressure, volume, and temperature

2nd Semester Chemistry Final Exam Study Guide

Brown University

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web determine the final temperature and state of the ammonia once the system reaches equilibrium assuming that no heat is lost to the surroundings use the information below normal boiling point of NH_3 33.3 °C ΔH_{vap} 23.4 kJ/mol specific heat NH_3 1.474 J/g·K specific heat NH_3 2.06 J/g·K

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web once all of the material has solidified at the time indicated by point c the cooling will continue at a rate determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase

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