

AP Biology Multiple Choice Exam Questions and Answers

A dog is following the scent of a jackrabbit. Which of the following accurately describes how the dog's brain integrates information for smell?

- A) Chemoreceptors in the brain send impulses for smell in the nasal cavity.
- B) Chemoreceptor cells in the nasal cavity sends impulses to the appropriate area of the brain.
- C) Chemoreceptors on epithelial cells of the tongue send hormones to the appropriate area of the brain.
- D) Receptors originating in the nose send action potentials to the motor regions of the brain. - Answer- B) Chemoreceptor cells in the nasal cavity sends impulses to the appropriate area of the brain.

Thrips are insects that feed on rose pollen. Scientists noted that the thrips population increased in the spring and decreased dramatically during the summer. The researchers hypothesized that food abundance was the limiting factor for the population. Which of the following types of data would be most useful for the scientists to collect at regular intervals on a designated test plot of rose plants?

- A) Amount of sunlight (hours/day)
- B) Mean temperature (°C)
- C) Density of rose pollen produced (g/m²)
- D) Amount of pollen produced by each flower (g/flower) - Answer- C) Density of rose pollen produced (g/m²)

If ATP breakdown (hydrolysis) is inhibited, which of the following types of movement across cell membranes is also inhibited?

- A) Movement of oxygen into a cell
- B) Movement of water through aquaporins
- C) Passage of a solute against its concentration gradient
- D) Facilitated diffusion of a permeable substance - Answer- C) Passage of a solute against its concentration

Undersea landslides can disrupt marine habitats by burying organisms that live on the ocean floor. The graph above shows the size of a population of a certain organism that lives on the ocean floor. The population was affected by a recent landslide at the time

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AW Rasmussen



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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(s) \rightleftharpoons \text{Ca}^{2+}(aq) + 2\text{OH}^-(aq)$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is

used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $Ca(OH)_2 \rightarrow Ca^{2+} + 2OH^-$ $5 \times 10^{-2} M$ calcium ions $5 \times 10^{-2} M$ $0 M$ hydroxide ions calculate calcium hydroxide K_{sp} using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K_{sp} of a slightly soluble basic salt and how solubility changes with temperature K_{sp} is another form of equilibrium expression that deals with solubility the reaction that will take place is $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$

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web question experiment 8 determining K_{sp} for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity $0.050 M$ volume of filtered $Ca(OH)_2$ solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH^- in saturated $Ca(OH)_2$ Ca^{2+} in saturated $Ca(OH)_2$ M K_{sp} of $Ca(OH)_2$ average K_{sp} of $Ca(OH)_2$

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web the chemical equation used was $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^-(aq)$ the solubility product was determined to be $K_{sp} Ca^{2+} OH^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as Ca_3PO_4 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} Ca_3PO_4 \rightleftharpoons 3Ca^{2+} + 4PO_4^{3-}$

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca(OH)_2 was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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