

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid or base solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the K_w value of 1.0×10^{-14} at 25°C. Assume that all solutions are aqueous.

1. What products are _____ when hydrochloric acid (HCl) ionizes in water? The concentration of these acids produced in 0.100 M hydrochloric acid is _____ M.
2. Hydrochloric acid ionizes in water _____ or _____ molecules _____ ion concentration is _____ M. The _____ ion concentration is _____ M.
3. _____ ionizes in water to form _____ and _____ ions. The _____ ion concentration is _____ M. The _____ ion concentration is _____ M.
4. _____ ionizes in water to form _____ and _____ ions. The _____ ion concentration is _____ M. The _____ ion concentration is _____ M.

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Christian Drosten



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