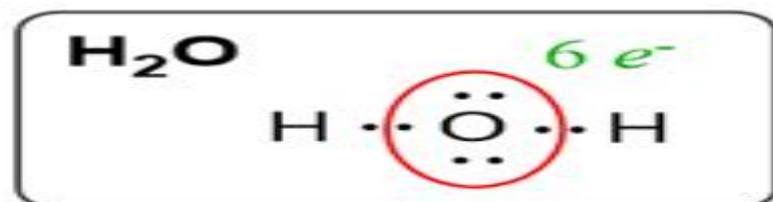


## Formal Charge

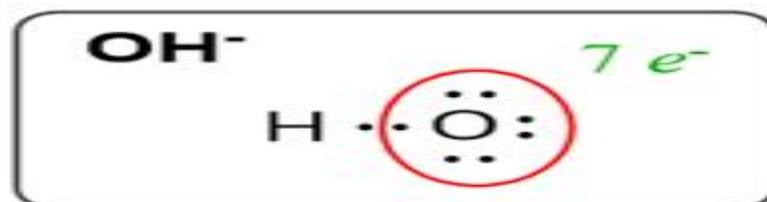
Oxygen is in group 6, so it has 6 valence  $e^-$ .



Valence  $e^-$  - 6

Assigned  $e^-$  - 6

Formal Charge = 0



Valence  $e^-$  - 6

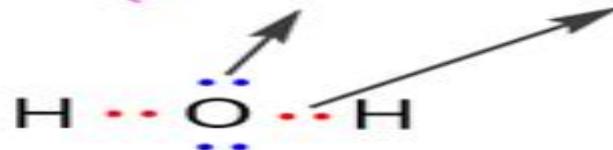
Assigned  $e^-$  - 7

Formal Charge = -1

*Formal charge is the difference between the number of valence electrons and the sum of the nonbinding and half of the bonding electrons.*

$$\text{FC} = \# \text{ valence } e^- - (\# \text{ nonbonding } e^- + 1/2 \# \text{ bonding } e^-)$$

**H<sub>2</sub>O**     $\text{FC} = 6 - (2 \times 2 + 0.5 \times 4) = 0$



**OH<sup>-</sup>**     $\text{FC} = 6 - (3 \times 2 + 0.5 \times 2) = -1$



# Formal Charges

**S Baum**



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