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Albert A Gayle



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mole moles of H_2O 40.1822 moles therefore according to the equation

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time the total pressure $p_a \times \text{total mole fraction}$ a left frac mole a moles total right sum $n_1 \text{ all } x_{n_1}$ mole fraction of substance
a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
mole fraction is not dependent on the temperature

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mass C_6H_6 128.17 g mol⁻¹ hence C_6H_6 s number of moles $\frac{30}{128.17} = 0.234$ mol molar mass of carbon tet CCl_4 153.84 g mol⁻¹

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