

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 5.5$       (C)  $\text{pH} = 6.5$       (D)  $\text{pH} = 7.5$       (E)  $\text{pH} = 8.5$

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 5.5$       (C)  $\text{pH} = 6.5$       (D)  $\text{pH} = 7.5$       (E)  $\text{pH} = 8.5$

3. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 5.5$       (C)  $\text{pH} = 6.5$       (D)  $\text{pH} = 7.5$       (E)  $\text{pH} = 8.5$

4. Which of the following is a weak acid? (correct)



5. A solution of an acid has a pH of 2.5. The concentration of the acid is 0.1 mol/L. What is the  $\text{pK}_a$  of the acid? (correct)

(A)  $\text{pK}_a = 2.5$       (B)  $\text{pK}_a = 3.5$       (C)  $\text{pK}_a = 4.5$       (D)  $\text{pK}_a = 5.5$       (E)  $\text{pK}_a = 6.5$

(A)  $\text{pK}_a = 2.5$       (B)  $\text{pK}_a = 3.5$       (C)  $\text{pK}_a = 4.5$       (D)  $\text{pK}_a = 5.5$       (E)  $\text{pK}_a = 6.5$

6. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A)  $\text{pH} = 4.5$       (B)  $\text{pH} = 5.5$       (C)  $\text{pH} = 6.5$       (D)  $\text{pH} = 7.5$       (E)  $\text{pH} = 8.5$

# Multiple Choice Stoichiometry Test

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membentuk perkerasan jalan sebagai sarana maupun prasarana yang dengan sni 03 1731 1989 atau cbr laboratorium sesuai dengan sni 03 1744 1989 apabila tanah dasar memiliki nilai cbr kurang dari 2 maka dianggap mempunyai nilai

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