

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) volume, molarity, pH, concentration of ions, amount of substance

(A)  $V = \text{dm}^3$ ,  $M = \text{mol dm}^{-3}$

(B)  $V = \text{dm}^3$ ,  $n = \text{mol}$ ,  $pH = -\log_{10} \frac{[H^+]}{1 \text{ mol dm}^{-3}}$ ,  $c = \text{mol dm}^{-3}$ ,  $[H^+] = \text{mol dm}^{-3}$   
 (C)  $V = \text{dm}^3$ ,  $n = \text{mol}$ ,  $pH = -\log_{10} \frac{[H^+]}{1 \text{ mol dm}^{-3}}$ ,  $c = \text{mol dm}^{-3}$ ,  $[H^+] = \text{mol dm}^{-3}$

2. An acid-base indicator has a  $pK_a$  of 4.5. In a solution of the indicator, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A)  $pH = 4.5 + \log_{10} 10$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

(B)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

3. An acid-base indicator has a  $pK_a$  of 4.5. In a solution of the indicator, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

4. Which indicator is most suitable for the titration of a weak acid with a strong base? (correct)

(A)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

(B)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

5. A solution of an indicator with a  $pK_a$  of 4.5 is titrated with a strong base. At the end of the titration, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

(B)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

6. An acid-base indicator has a  $pK_a$  of 4.5. In a solution of the indicator, the ratio of the concentrations of the two forms is 10:1. What is the pH of the solution? (correct)

(A)  $pH = 4.5$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$ ,  $[OH^-] = 10^{-9.5} \text{ mol dm}^{-3}$ ,  $[H^+] = 10^{-4.5} \text{ mol dm}^{-3}$

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