

**KEY**

1 26. Which is the stronger base,  $\text{HCO}_3^-$  or  $\text{HCOO}^-$ ? (1 mark)  $\text{HCO}_3^-$

27. Classify each of the following as: a strong acid (SA), weak acid (WA), strong base (SB), weak base (WB) or a spectator ion (S). (10 marks)

10

- |                                |           |                             |           |
|--------------------------------|-----------|-----------------------------|-----------|
| a) $\text{F}^-$                | <u>WB</u> | f) $\text{Cl}^-$            | <u>S</u>  |
| b) $\text{HIO}_3$              | <u>WA</u> | g) $\text{NH}_3$            | <u>WB</u> |
| c) $\text{NO}_3^-$             | <u>S</u>  | h) $\text{O}^{2-}$          | <u>SB</u> |
| d) $\text{HClO}_4$             | <u>SA</u> | i) $\text{CH}_3\text{COOH}$ | <u>WA</u> |
| e) $\text{C}_2\text{O}_4^{2-}$ | <u>WB</u> | j) $\text{ClO}_4^-$         | <u>S</u>  |

28. What is the  $[\text{OH}^-]$  in a solution made by adding 0.060 moles of calcium oxide to 500.0 mL water? Be careful! (2 marks)

2

$$M = \frac{\text{mol}}{L} = \frac{0.060 \text{ mol CaO}}{0.500 \text{ L}} = 0.12 \text{ M} = [\text{CaO}]$$

$$0.12 \text{ CaO} \rightarrow \text{Ca}^{2+} + 0.12 \text{ O}^{2-} \Rightarrow \text{O}^{2-} + \text{H}_2\text{O} \rightarrow 2\text{OH}^-$$

Answer  $[\text{OH}^-] = 0.24 \text{ M}$

29. What is the  $[\text{H}_3\text{O}^+]$  in a solution made by adding 0.020 moles of nitric acid to 500.0 mL of water? (2 marks)

2

$$[\text{HNO}_3] = \frac{0.020 \text{ mol}}{0.500 \text{ L}} = 0.040 \text{ M}$$

SA so  $[\text{H}_3\text{O}^+] = [\text{acid}] = 0.040 \text{ M}$

Answer  $[\text{H}_3\text{O}^+] = 0.040 \text{ M}$

30. If 0.10 M  $\text{HSO}_4^-$  is mixed with 0.10 M  $\text{HC}_2\text{O}_4^-$ , which species will donate a proton? (1 mark)

1

Answer  $\text{HC}_2\text{O}_4^-$

31. If 0.10 M  $\text{HSO}_4^-$  is mixed with 0.10 M  $\text{HC}_6\text{H}_5\text{O}_2^-$ , which species will donate a proton? (1 mark)

1

Answer  $\text{HSO}_4^-$

$\frac{17}{17}$

# 19 3 Acids And Bases Answer Key Chemistry

**SJ Ball**



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