

18.4 Molecular Structure: Mass

Multiple Choice

1. The molar masses of quantities in chemical equations are based on _____
 (A) percent composition (B) percent yield
 (C) molar mass (D) stoichiometric coefficients
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____ of the total mass of the compound.
 (A) 15.0% (B) 13.3%

3. In a gas the other reactants represent oxygen and sulfur dioxide, 2.5 g of gas was produced by the reaction, and only 0.5 g of gas, identify the percent yield of sulfur in this reaction?



- (A) 20% (B) 25% (C) 20.5% (D) 20.1% (E) 20.2%

4. In this reaction $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$, which of the following would be correct for the reaction of 100g of CO and 100g of O_2 ?

- (A) 100g (B) 150g (C) 200g (D) 250g

5. Which of the following is correct for the following equation?



- (A) 1 (B) 2 (C) 3 (D) 4

6. Aluminum reacts with HCl to produce hydrogen gas and aluminum chloride. If 2.50 g of aluminum reacts with 10.00 g of HCl, what mass of H_2 gas is produced?



- (A) 0.889 g (B) 1.50 g (C) 2.50 g (D) 3.00 g (E) 5.50 g

7. In a full combustion, structure-molecular weight $\text{C}_2\text{H}_6\text{O}$ and other gases and CO_2 and H_2O are produced. After the reaction was completed, 100.0 g of CO_2 was produced for the reaction. What was the limiting reactant?



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