

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (E) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (F) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

2. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. The pH of the solution is (correct)
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (E) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

3. The pH of a 0.1 M acetic acid solution is 3.0. The concentration of acetate ions is (correct)
- (A) 0.01 M (B) 0.02 M (C) 0.03 M (D) 0.04 M (E) 0.05 M

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

4. Which of the following is a weak acid? (correct)
- (A) HCl (B) H_2SO_4 (C) H_2CO_3 (D) HNO_3

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

5. A 0.1 M acetic acid solution is mixed with 0.1 M sodium acetate. The pH of the solution is (correct)
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(E) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (F) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (G) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (H) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

6. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. The pH of the solution is (correct)
- (A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0

(A) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (B) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (C) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$ (D) $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

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Peipei Pang



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