

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) volume, molarity, pH, concentration of ions, amount of substance

(A) $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$

(B) $p = 10^{-\text{pH}}$, $V = \text{dm}^3$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$
 (C) $V = \text{dm}^3$, $p = 10^{-\text{pH}}$, $M = \text{mol dm}^{-3}$, $n = \text{mol}$ (not all the same)

2. An acid-base indicator has a $\text{p}K_{\text{a}}$ of 4.5. The indicator is colourless in its protonated form and pink in its deprotonated form. What colour will the indicator be in a solution of pH 4.5? (correct)

(A) $\text{pH} < \text{p}K_{\text{a}}$, colourless
 (B) $\text{pH} > \text{p}K_{\text{a}}$, pink

(C) $\text{pH} = \text{p}K_{\text{a}}$, pink
 (D) $\text{pH} = \text{p}K_{\text{a}}$, colourless
 (E) $\text{pH} = \text{p}K_{\text{a}}$, pink

3. An acid-base indicator has a $\text{p}K_{\text{a}}$ of 4.5. The indicator is colourless in its protonated form and pink in its deprotonated form. What colour will the indicator be in a solution of pH 5.5? (correct)

(A) $\text{pH} < \text{p}K_{\text{a}}$, colourless
 (B) $\text{pH} < \text{p}K_{\text{a}}$, pink
 (C) $\text{pH} > \text{p}K_{\text{a}}$, colourless
 (D) $\text{pH} > \text{p}K_{\text{a}}$, pink

4. Which indicator is most suitable for the titration of a weak acid with a strong base? (correct)

(A) $\text{p}K_{\text{a}} = 3.5$, $\text{pH} = 3.5$
 (B) $\text{p}K_{\text{a}} = 7.5$, $\text{pH} = 7.5$

(C) $\text{p}K_{\text{a}} = 3.5$, $\text{pH} = 7.5$
 (D) $\text{p}K_{\text{a}} = 7.5$, $\text{pH} = 3.5$

5. A solution of sodium hydroxide and a solution of acetic acid are mixed together to form a buffer solution. If a small amount of acetic acid is added to the buffer, what will happen to the pH? (correct)



(E) pH will decrease
 (F) pH will increase
 (G) pH will remain the same
 (H) pH will decrease

6. An acid-base indicator, structure (I), has a $\text{p}K_{\text{a}}$ of 4.5. The indicator is colourless in its protonated form and pink in its deprotonated form. What colour will the indicator be in a solution of pH 4.5? (correct)



Multiple Choice Stoichiometry Test

Puja Mehta

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web ufugaji kuku kiholela haufai mkm kwenye mtandao kuku kama ilivyo kwa mifugo wengine wanahitaji kuwa na banda na kupatiwa matunzo sahihi kwa kufanya hivyo utaweza