



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>hydrogen ion</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li><li>pH</li></ul>	<ul style="list-style-type: none"><li>hydrogen ion concentration</li><li>acidic</li><li>basic</li><li>neutral</li><li>titration</li><li>titration curve</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = 7$  at 25°C

### Principles of Acids and Bases

Acids and bases are substances that have the ability to change the properties of other substances.

Acids are substances that contain hydrogen ions. The concentration of these ions determines the strength of the acid.

Strong acids contain strong acids. Strong acids dissociate completely. Weak acids dissociate partially.

Acids are substances that have the ability to change the properties of other substances.

<ul style="list-style-type: none"><li>strong acid</li><li>weak acid</li><li>acidic</li><li>basic</li><li>neutral</li><li>titration</li><li>titration curve</li></ul>	<ul style="list-style-type: none"><li>hydrogen ion</li><li>hydrogen ion concentration</li><li>acid</li><li>base</li><li>strong acid</li><li>weak acid</li><li>pH</li></ul>
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## 19 2 Hydrogen Ions And Acidity Section Review

**Robert R. Redfield**

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