

1

26. Which is the stronger base, HCO_3^- or HCOO^- ? (1 mark) HCO_3^-

27.

27. Classify each of the following as: a strong acid (SA), weak acid (WA), strong base (SB), weak base (WB) or a spectator ion (S). (10 marks)

a) F^-

WB

f)

 Cl^-

S

b) HIO_3

WA

g)

 NH_3

WB

c) NO_3^-

S

h)

 O^{2-}

SB

d) HClO_4

SA

i)

 CH_3COOH

WA

e) $\text{C}_2\text{O}_4^{2-}$

WB

j)

 ClO_4^-

S

28.

What is the $[\text{OH}^-]$ in a solution made by adding 0.060 moles of calcium oxide to 500.0 mL water? Be careful! (2 marks)

$$\frac{\text{M} \cdot \text{mol}}{\text{L}} = \frac{0.060 \text{ mol CaO}}{0.5000 \text{ L}} = 0.12 \text{ M} = [\text{CaO}]$$



29.

What is the $[\text{H}_3\text{O}^+]$ in a solution made by adding 0.020 moles of nitric acid to 500.0 mL of water? (2 marks)

$$[\text{HNO}_3] = \frac{0.020 \text{ mol}}{0.500 \text{ L}} = 0.040 \text{ M} \quad \text{SA so } [\text{H}_3\text{O}^+] = [\text{acid}] = 0.040 \text{ M}$$

Answer $[\text{H}_3\text{O}^+] = 0.040 \text{ M}$

30.

If 0.10 M HSO_4^- is mixed with 0.10 M HC_2O_4^- , which species will donate a proton? (1 mark)

Answer

 HC_2O_4^-

31.

If 0.10 M HSO_4^- is mixed with 0.10 M $\text{HC}_2\text{H}_3\text{O}_2^-$, which species will donate a proton? (1 mark)

Answer

 HSO_4^-

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17

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