

18.4 Molecular Structure: Mass

Multiple Choice:

1. The molar masses of quantities in chemical equations are based _____
 (A) percent composition (B) percent yield
 (C) molecular weight (D) stoichiometric coefficients
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is approximately the same as percent of _____?



3. In a gas the other reactions represent oxygen mixed with other elements. If 1.1 g of oxygen combined to form formaldehyde, and only 0.6 g of oxygen remained, what percent of oxygen is in the reaction?



4. In the reaction $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{C}_2\text{H}_4\text{O}_2$, what is the percent yield of oxygen?
 (A) 100% (B) 100% (C) 100% (D) 100% (E) 100%

5. Percent of carbon in C_2H_6 is approximately the same as percent of _____?



6. A mixture of 10 g of hydrogen and 10 g of oxygen reacts to produce water. If 1.1 g of hydrogen is left and 10.9 g of water is produced, what mass of O_2 was produced?



7. In a full combustion, structure $\text{C}_2\text{H}_4\text{O}_2$ reacts with O_2 to form CO_2 and H_2O . What percent of O_2 is in the reaction? (A) 100% (B) 100% (C) 100% (D) 100% (E) 100%



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N Colangelo

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