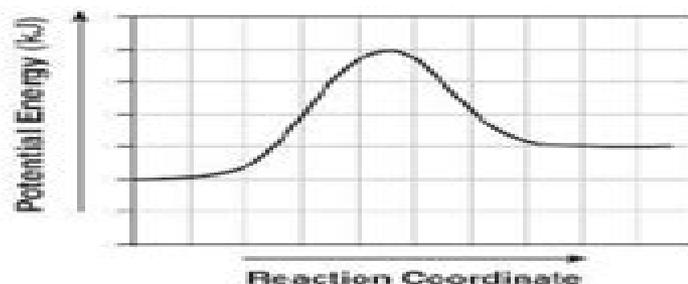


- 1) When properly balanced, the coefficient for hydrochloric acid (HCl) would be?
 $\underline{\quad} \text{Al} + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{AlCl}_3 + \underline{\quad} \text{H}_2$
- 2) The potential energy diagram for a chemical reaction is shown below



Each interval on the axis labeled "Potential Energy (kJ)" represents 40 kilojoules. What is the heat of reaction (ΔH)?

- 3) The general equation shown is a reaction that represents what type of reaction? _____ is this reaction endothermic or exothermic? (circle one)



- 4) Write a balanced equation for the word equation shown below:
 Aluminum + Sulfuric acid \rightarrow Aluminum sulfate + hydrogen gas

- 5) A process that absorbs heat is: endothermic or exothermic (circle one) Would the products have more energy or less energy than the reactants?
- 6) The two beakers of boiling water shown below are on hot plates set at different settings. However, the water in both beakers is at the same temperature. How is the incoming energy from the hot plates being used by the water if it is not increasing the temperature?



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S Ashworth



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1. Understanding the eBook 2nd Semester Final Exam Review Science Answer
 - The Rise of Digital Reading 2nd Semester Final Exam Review Science Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Final Exam Review Science Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Final Exam Review Science Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Final Exam Review Science Answer
 - Personalized Recommendations
 - 2nd Semester Final Exam Review Science Answer User Reviews and Ratings
 - 2nd Semester Final Exam Review Science Answer and Bestseller Lists
5. Accessing 2nd Semester Final Exam Review Science Answer Free and Paid eBooks

- 2nd Semester Final Exam Review Science Answer Public Domain eBooks
 - 2nd Semester Final Exam Review Science Answer eBook Subscription Services
 - 2nd Semester Final Exam Review Science Answer Budget-Friendly Options
6. Navigating 2nd Semester Final Exam Review Science Answer eBook Formats
- ePub, PDF, MOBI, and More
 - 2nd Semester Final Exam Review Science Answer Compatibility with Devices
 - 2nd Semester Final Exam Review Science Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 2nd Semester Final Exam Review Science Answer
 - Highlighting and Note-Taking 2nd Semester Final Exam Review Science Answer
 - Interactive Elements 2nd Semester Final Exam Review Science Answer
8. Staying Engaged with 2nd Semester Final Exam Review Science Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Final Exam Review Science Answer
9. Balancing eBooks and Physical Books 2nd Semester Final Exam Review Science Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Final Exam Review Science Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Final Exam Review Science Answer
- Setting Reading Goals 2nd Semester Final Exam Review Science Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Final Exam Review Science Answer
- Fact-Checking eBook Content of 2nd Semester Final Exam Review Science Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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