

Name Key

1. Hydrogen, H<sub>2</sub>, is used as rocket fuel. The hydrogen is burned in oxygen, O<sub>2</sub>, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H<sub>2</sub> is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?

reverse so change sign of ΔH = 1077 kJ

$$\frac{1}{2} \text{CO}_2(\text{g}) + \text{SO}_2(\text{g}) \rightarrow \frac{1}{2} \text{CS}_2(\text{l}) + \frac{3}{2} \text{O}_2(\text{g})$$

$$\frac{+1077 \text{ kJ}}{2} = 538.5 \text{ kJ}$$

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N<sub>2</sub>H<sub>4</sub>, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Reverse  
doubly quantity

$$\begin{array}{r} \text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow \text{N}_2\text{H}_4(\text{l}) + \text{O}_2(\text{g}) \quad \Delta H = 622.2 \text{ kJ} \\ 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = (2)(-285.8) \text{ kJ} \\ \hline \text{N}_2(\text{g}) + 2 \text{H}_2(\text{g}) \rightarrow \text{N}_2\text{H}_4(\text{l}) \quad \Delta H = +50.6 \text{ kJ} \end{array}$$

# 19 3 Chemistry Practice Problems Answer Key

**Padhraic Smyth**



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## **Table of Contents 19 3 Chemistry Practice Problems Answer Key**

1. Understanding the eBook 19 3 Chemistry Practice Problems Answer Key
  - The Rise of Digital Reading 19 3 Chemistry Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Chemistry Practice Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Chemistry Practice Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Chemistry Practice Problems Answer Key
  - Personalized Recommendations
  - 19 3 Chemistry Practice Problems Answer Key User Reviews and Ratings
  - 19 3 Chemistry Practice Problems Answer Key and Bestseller Lists

5. Accessing 19 3 Chemistry Practice Problems Answer Key Free and Paid eBooks
  - 19 3 Chemistry Practice Problems Answer Key Public Domain eBooks
  - 19 3 Chemistry Practice Problems Answer Key eBook Subscription Services
  - 19 3 Chemistry Practice Problems Answer Key Budget-Friendly Options
6. Navigating 19 3 Chemistry Practice Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 3 Chemistry Practice Problems Answer Key Compatibility with Devices
  - 19 3 Chemistry Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 3 Chemistry Practice Problems Answer Key
  - Highlighting and Note-Taking 19 3 Chemistry Practice Problems Answer Key
  - Interactive Elements 19 3 Chemistry Practice Problems Answer Key
8. Staying Engaged with 19 3 Chemistry Practice Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 3 Chemistry Practice Problems Answer Key
9. Balancing eBooks and Physical Books 19 3 Chemistry Practice Problems Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 3 Chemistry Practice Problems Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Chemistry Practice Problems Answer Key
  - Setting Reading Goals 19 3 Chemistry Practice Problems Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Chemistry Practice Problems Answer Key
  - Fact-Checking eBook Content of 19 3 Chemistry Practice Problems Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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