



EXPLAINING MUSIC

ESSAYS AND EXPLORATIONS

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LEONARD B. MEYER

Explaining Music Essays And Explorations

JG Myers



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web mar 7 2021 unlike strong acids or bases the shape of the titration curve for a weak acid or base depends on the pK_a or pK_b of the weak acid or base being titrated a solution ph as a function of the volume of 1.00 M NaOH added to 10.00 mL of 1.00 M solutions of weak acids with the indicated pK_a values

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17 3 2 1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = \log 0.100 = 1.000$ $\text{ph} = \log 0.100 = 1.000$ b titrant volume 12.50 ml

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web at the equivalence point in an acid base titration moles of base = moles of acid and the solution only contains salt and water acid base titrations are monitored by the change of ph as titration progresses

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titration curve is the plot of the ph value of a solution vs the amount of titrant added

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web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there is only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow

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web now realizing that heating and cooling curves represent the amount of heat absorbed or released by a substance during phase changes remember heat uses the variable q here we have a heating curve versus a cooling curve in a heating curve our substance is absorbing heat

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web describe the processes represented by typical heating and cooling curves and compute heat flows and enthalpy changes accompanying these processes explain the construction and use of a typical phase diagram

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web 11 12 heating and cooling curves when a substance isolated from its environment is subjected to heat changes corresponding changes in temperature and phase of the substance is observed this is graphically represented by heating and cooling curves

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web determine the final temperature and state of the ammonia once the system reaches equilibrium assuming that no heat is lost to the surroundings use the information below normal boiling point of NH_3 33.3 $^\circ\text{C}$ ΔH_{vap} 23.4 kJ/mol specific heat NH_3 1.474 J/g $^\circ\text{C}$ specific heat NH_3 2.06 J/g $^\circ\text{C}$

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web apr 3 2023 the experiment described above can be summarized in a graph called a heating curve figure below figure 11.16.1 in the heating curve of water the temperature is shown as heat is continually added changes of state occur during plateaus because the temperature is constant

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web mar 29 2013 heating curves imagine that you have a block of ice that is at a temperature of 30 $^\circ\text{C}$ well below its melting point the ice is in a closed container as heat is steadily added to the ice block the water molecules will begin to vibrate faster and faster as they absorb kinetic energy

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web the cooling curve a plot of temperature versus cooling time in figure pageindex 4 plots temperature versus time as a 75 g sample of steam initially at 1 atm and 200 $^\circ\text{C}$ is cooled although we might expect the cooling curve to be the mirror image of the heating curve in figure pageindex 3 the cooling curve is not an

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web once all of the material has solidified at the time indicated by point c the cooling will continue at a rate determined by the heat capacities of the two solids a and b the composition and of course the geometry of the experimental set up by measuring cooling curves for samples of varying composition one can map the entire phase

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web for water this temperature is 100 c because the boiling point for water is 100 c different substances have different melting points and boiling points but the shapes of their heating curves are very similar for example this is the heating curve for iron a metal that melts at 1538 c and boils at 2861 c cooling curves

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