

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance** Basic SI Unit? **kg**
 - Volume of an object? **ruler or water displacement method** Basic SI Unit? **cm³ or mL**
 - Length of an object? **Ruler** Basic SI Unit? **meter**
 - Density of an object? **graduated cylinder** Basic SI Unit? **mL**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 33 g and has a volume of 3 cm³? **10g/cm³**
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. **solid H₂O less dense than liquid H₂O**

2. Chemistry Basics

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move fixed orbits around the atom's nucleus in specific layers, or electron shells . Atoms absorb or give off energy when the electrons move from one shell (energy level) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

4. Periodic Table

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
 - The atomic number represents the number of **protons**.
 - The atomic mass represents the number of **protons and neutrons**.

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Enjoying the Beat of Expression: An Emotional Symphony within **2nd Semester Final Exam Study Guide Answers**

In a global eaten by monitors and the ceaseless chatter of instant connection, the melodic splendor and psychological symphony developed by the published term frequently disappear in to the background, eclipsed by the persistent noise and disruptions that permeate our lives. But, situated within the pages of **2nd Semester Final Exam Study Guide Answers** a charming fictional value overflowing with natural thoughts, lies an immersive symphony waiting to be embraced. Constructed by a wonderful composer of language, that fascinating masterpiece conducts visitors on an emotional journey, well unraveling the hidden melodies and profound affect resonating within each cautiously crafted phrase. Within the depths with this moving analysis, we will discover the book is central harmonies, analyze its enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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